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Grams and liters molarity calculations worksheet

Grams and Liters Of Molarity Calculations Sheet Name _____ Period _____ Molyaricity is the number of moles dissolved in one liter of solution. Units, therefore, there are moles per liter, in particular, it moles dissolve per liter of solution. Instead of writing moles per liter, these units are abbreviated capitol M. Example 1: What is the mole of a 3L solution containing 5.00 NaOH moles? M = Moles litre of solution M = 5,00 moles 3 litres solution M = 1,67 moles/litre Example 2: What is the mole of a 10L solution containing 5,00 moles Na2CO3? M = Moles liter of solution M = 5,00 moles 10 liters solution M = .5 moles / liter Problems: Solution description Molilarity: 1. 4,67 moles Li2SO3 dissolved to 2,04 liters of solution. 2. 0.629 moles Al2O3 to make 1.500 liters of solution. 3. 8,00 moles EDTA to 7,56 liters of solution We can also determine the moles dissolved in the solution, if we know the volume and mole. Example 3: How many Na2CO3 moles are in 10.0 l of 2.0 M solution? # moles = liter of solution x Molecy = 10.0 L x 2.0 M # moles = 20 Problems: Solution description # Moles: 1. How many moles Cu(NO3)2 are in 3.35 liters 2.00 M Cu (NO3)2 solution? 2. How many MgCO3 moles are in 3.00 litres of MgCO3 1.50 M solution? 3. How many NaCl moles are in 100,0 ml of 0,20 M solution? 4. How many Cu(NO3)2 moles are in 2.35 litres of 2.00 M solution? 5. How many pb(NO3)2 moles are in 160,0 ml of molar solution 0,415? 6. How many MgCO3 moles is in 3,00 l of 0,500 M solution? 7. How many Na2O moles are in 6,20 l of molar solution 3,76 molar? We can also determine the number of liters of solution to use if we know the number of moles needed and the molarity of the solution Example 4: What volume (in ml) 18.0 M H2SO4 is needed to accommodate 2.45 moles H2SO4? liters of solution = mole dissolved mole mole mole liters solution = 2,45moles 18,0 M liters of solution = .136 = 136 ml Problems: Description of the solution # Liters: 1. What will be the final volume of solution containing 4,67 moles Li2SO3 dissolved to solve 3,89 M? 2. What will be the final volume of solution containing 4,907 Al2O3 moles to make a 0,500 M solution? We can replace moles in the equation and use the following: Liters of solution = grams of soluble molar mass Molarity Example 5: What will be the final volume of solution containing 45,7 g NaOH to make a solution .5 M? liters of solution = 45.7 g 39.99 g/mole 0.50 M liters of solution = 2.29 L Problems: Solution description # Liters: 1. What will be the final volume of solution containing 0.783 grams of Na2CO3 to make a solution 34.8 M? 2. What will be the final volume of solution containing 8,97 grams (NH4)2CO3 to make a molar solution of 0,250? 3. What will be the final volume of solution containing 48.00 pbCl2 to form a molar solution 5.0? We can also determine the molalarity of the solution if we know the number of grams and the volume of the solution. M = grams solution of molar solution solution example 6: What is the molarity of a 3L solution containing 5,00 g NaOH? M = grams molar solution Solution M = 5,00 g 39,99 g/mol 3 liters solution M = .042 moles/liter Problems: Solution Description Molarity: 1. Seawater contains about 28,0 g NaCl per litre. What is the molalarity of sodium chloride in seawater? 2. What is the molocity of a solution containing 5,30 g Of Na2CO3 dissolved in a 400,0 ml solution? 3. What is the molocity of a solution containing 5,00 g of NaOH dissolved in 750,0 ml of solution? 4. What is the molocity of a solution containing 4,783 grams of Na2CO3 dissolved in 1,00 litres of solution? 5. What is the molocity of a solution containing 8,97 grams (NH4)2CO3 dissolved in 250 ml of solution? 6. What is the molocity of a solution containing 0,348 grams of PbCl2 dissolved in 45,0 ml of solution? Changing the layout of the above equation, we can easily calculate the amount of chemical needed in grams for the solution of a given final volume and molarity. we play the chemical = M x molar mass x liter of solution It should be noted that the units cancel and you remain with grams. Example 7: How many grams of Cu(NO3)2 are needed to make 0.289 liters of 0.0300 M Cu(NO3)2 solution? Grams = M x molar mass x litre of solution Gram = 0,0300 moles/litre x 187.57g/mole x 0.289 litres Gram = 1.63 Problems: Solution description We play 1. How many grams of Pb(NO3)2 are needed for 16.00 millilitres of 5.90 molar Solution Pb(NO3)2? 2. How many grams of NaF is needed to make 508 ml of 2.75 Mol NaF solution? 3. How many grams of Na2O is needed to make 6,20 l of Molar solution 3,76 Mola Na2O? 4. How many grams of KCl are needed for a 0.500 L solution of 1.00 M KCl? 5. How many grams of CaCl2 is needed to make 4.35 l of 3.50 M CaCl2 solution? 6. How many grams of Ca(OH)2 is needed to make 100.0 ml of 0.250 M solution? 7. What weight (in grams) of KCl is there in a 2.50 liter 0.50 M KCl solution? We play gram/moles calculations: Given the following, find the number of grams. (Molar molars): Problem: We play: Problem: 1) 4 moles Cu(CN)2 5) 9,3 x 10-3 moles SmO 2) 5,6 moles C6H6 6) 6,6 moles ZnO 3) 21,3 moles BaCO3 7) 5.6 moles 4 moles K2SO4 4) 1.2 moles (NH4)3PO3 8) 88.4 moles NI3 Determine the dissolution grams to prepare these solutions: Problem: We play: Problem: We play: a) 0.289 liters 0.00300 M Cu(NO3)2 solution. (d) 6,20 l of 3,76 molar Na2O solution. (b) 16,00 millilitres of 5,90 molar solution Pb(NO3)2. (e) 0,500 l solution of 1,00 M KCl.c) 508 ml of 2,75 molar NaF solution. (f) 4,35 l of 3,50 M CaCl2 solution. Determine the final volume of these solutions: Problem: Liters: Problem: Liters: a) 4.67 moles Li2SO3 dissolved to make a solution of 3.89 M. d) 8.97 grams to make a 0,250 molar solution. (b) 4,907 Al2O3 moles to a solution of 0.500 M. e) 48,00 grams of PbCl2 to form a 5.0-molar solution. c) 0.783 grams of Na2CO3 to make a solution of 0.348 M. Summary of problems: Use what you know about molar problems that we have come up with so far to answer the following problems: Problem: Answer: 1) Find the number of moles dissolved 2.5 l ag2SO4 solution 4.6 M. 2) Find the number of moles in 0,5 l of NaOH solution 2,1 M. 3) Find the volume of 5,3 moles Of NaCl dissolved to make a solution of 0,51 M. 4) What is the volume (in liters) of 4,67 moles of KCl dissolved to make a solution of 3,7 M? 5) Find the molarity given 3.48 moles of CH3COOH at 14.3 L. 6) Given 22.98 g of HBr in 8.4 l, what is molarity? 7) 247.3 g C6H12O6 9.9 L Using this information, find the molalarity. 8) Znajdź liczbę gramów potrzebnych do wykonania 3,08 l roztworu KCN o masie 0,27 M. Grams/Moles Calculations Answer Key Given the following, find the number of moles: 1) 30 grams of H3PO4 0.31 moles 2) 25 grams of HF 1.25 moles 3) 110 grams of NaHCO3 1.31 moles 4) 1.1 grams of FeCl3 0.0068 moles 5) 987 grams of Ra(OH)2 3.80 moles 6) 564 grams of copper 0.11 moles 7) 12.3 grams of CO2 0.28 moles 8) 89 grams of Pb(CH3COO)4 0.20 moles Given the following, find the number of grams: 9) 4 moles of Cu(CN)2 462 grams 10) 5.6 moles of C6H6 436.8 grams 11) 21.3 moles of BaCO3 4202.5 grams 12) 1.2 moles of (NH4)3PO3 159.6 grams 13) 9.3 x 10-3 moles of SmO 1.5 grams 14) 6.6 moles of ZnO 537.2 grams 15) 5.4 moles of K2SO4 941.2 grams 16) 88.4 moles of NI3 34679.3 grams Деякі функції Word не відображаються в Google Документах. Якщо ви внесете зміни, налаштування цих функцій буде скинуто.Докладніше Gramy i litry Molaryczność Obliczenia Nazwa arkusza _Докладніше Gramy i litry Molaryczności Obliczenia Nazwa arkusza _Докладніше Gramy i litry Molary obliczenia Nazwa arkusza _Докладніше Gramy i litry Molaryczności Obliczenia Nazwa arkusza _Докладніше Gramy i litry Molaryścia Obliczenia Nazwa arkusza _Докладніше Gramy i litry Molaryczności Obliczenia Nazwa arkusza _Докладніше Gramy i litry Molaryczność Obliczenia Nazwa arkusza _Докладніше Gramy i litry Molaryczność Obliczenia Nazwa arkusza _До

Jednostki, dlatego są moli na litr, w szczególności to krety rozpuszczenie na litr roztworu. Instead of writing moles per liter, these units are abbreviated capitol M. Example 1: What is the mole of a 3L solution containing 5.00 NaOH moles? M = Moles litre of solution M = 5,00 moles 3 litres solution M = 1,67 moles/litre Example 2: What is the mole of a 10L solution containing 5,00 moles Na2CO3? M = Moles liter solution M = 5,00 moles 10 liters solution M = .5 moles / liter Problems: Solution description Mole: 4.67 moles li2SO3 dissolved to make 2.04 liters of solution. 2. 0.629 moles Al2O3 to make 1.500 liters of solution. 3. 8,00 moles EDTA to 7,56 liters of solution We can also determine the moles dissolved in the solution, if we know the volume and mole. Example 3: How many Na2CO3 moles are in 10.0 l of 2.0 M solution? # moles = liter solution x Molacyence # moles = 10.0 L x 2.0 M # moles = 20 Problems: Solution description # Moles: 1. 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